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ACIDS, BASES AND SALTS

GENERAL SCIENCE

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ACIDS

Introduction

- An acid is the compound which is capable of forming hydrogen ions (H^+) in aqueous solution whereas a base is the compound that forms hydroxyl ions (OH^-) in solution.

Acid

- The word 'acid' is derived from the Latin name "acidus" which means sour taste. Substances with sour taste are called acids.

Classification of Acids

Based on their sources:

- * Organic acids
- * Inorganic acids

Based on their Basicity

- * Monobasic Acid
- * Dibasic Acid
- * Tribasic Acid

Based on Ionisation

- * Strong Acids
- * Weak Acids

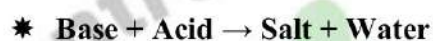
Based on Concentration

- * Concentrated Acid
- * Dilute Acid

BASES

Introduction

- According to Arrhenius theory, bases are substances that ionise in water to form hydroxyl ions (OH^-).
- There are some metal oxides which give salt and water on reaction with acids.
- A base reacts with an acid to give salt and water only.



Properties of Bases:

- They have bitter taste.
- Their aqueous solutions have soapy touch.
- They turn red litmus blue

pH Scale

- A scale for measuring hydrogen ion concentration in a solution is called pH scale.
- The 'p' in pH stands for '**potenz**' in German meaning power.

SALT

Introduction

- Salts are the products of the reaction between acids and bases.
- Salts produce positive ions and negative ions when dissolved in water.
- $\text{Acid} + \text{Base} \rightarrow \text{Salt} + \text{Water}$

Types of Salts

- (i) Normal Salts
- (ii) Acid Salts
- (iii) Basic Salts
- (iv) Double Salts

Properties of Salts

- ★ Salts are mostly solids which melt as well as boil at high temperature.
- ★ Most of the salts are soluble in water. For example, chloride salts of potassium and sodium are soluble in water. But silver chloride is insoluble in water.

Uses of Salts

- Common Salt (NaCl)
- Washing Soda (Sodium Carbonate-)
- Baking Soda (Sodium bicarbonate - NaHCO_3)